

Cornell Notes

Period 1

Name Ashley Martinez

Date Oct 27, 2011

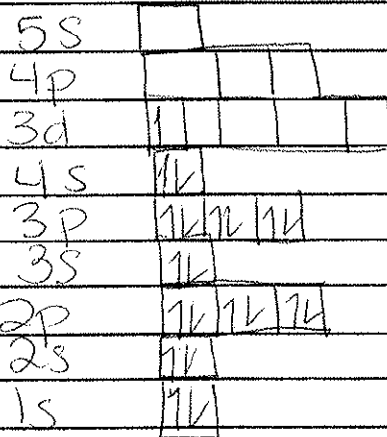
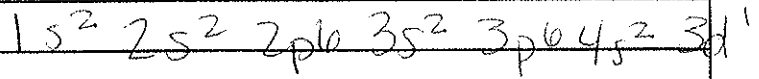
Topic

Class/Subject Chem / Fraguero

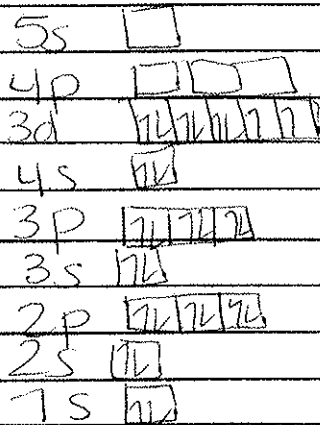
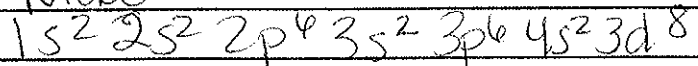
Write electron configuration + orbital diagram for:

7:55 am

1) Sodium



Nickel



* Study for Quiz tomorrow

— memorize everything that was labeled

on periodic table: family name, block which element is located, based on group location — how to determine charge of ion.

Cornell Notes

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D-block elements

- The middle of periodic table
- Transition metals

- Elements wide.
 - d-orbitals hold up to 10-electrons

S-block elements

- Group 1 (alkali metals) and 2 (alkaline earth metals)

- Outermost energy orbital:
 - * s-orbital

- Group 1
 - 1 valence electron
 - Gives up electron in chemical ions
 - Forms $1+$ ion

- Group 2
 - 2 valence electrons
 - Gives up 2 electrons
 - Forms $2+$ ion

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8:15
am

Aluminum

Boron

Nitrogen

Aluminum
 $1s^2 2s^2 2p^6 3s^2 3p^1$

3p	↑		
3s	↑↓		
2p	↑↓	↑↓	↑↓
2s	↑↓		
1s	↑↓		

Boron
 $1s^2 2s^2 2p^1$

2p	↑		
2s	↑↓		
1s	↑↓		

Nitrogen
 $1s^2 2s^2 2p^3$

2p	↑	↑	↑
2s	↑↓		
1s	↑↓		

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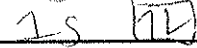
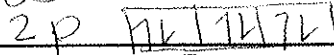
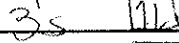
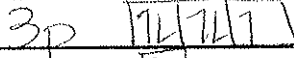
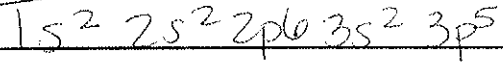
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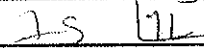
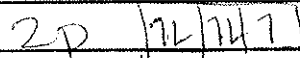
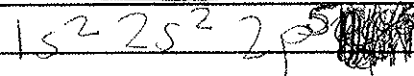
Chlorine

Chlorine



Fluorine

Fluorine



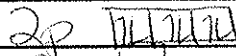
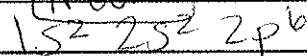
Helium

Helium



Neon

Neon



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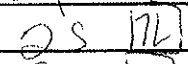
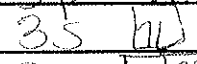
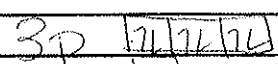
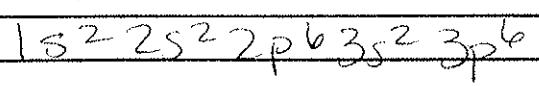
Topic

Class/Subject chem / Frago

8:30 am

Argon

Argon



P-block elements

- 6 elements wide
- p-orbitals hold up to 10 electrons

- Elements in the same column (group) will have the same # of valence electrons in the outermost p-orbital.

Electron Configuration $\frac{1}{3}$
The Periodic Table

- Groups in Periodic Table - Elements found in the same group have similar properties because of similar valence electron configuration.

