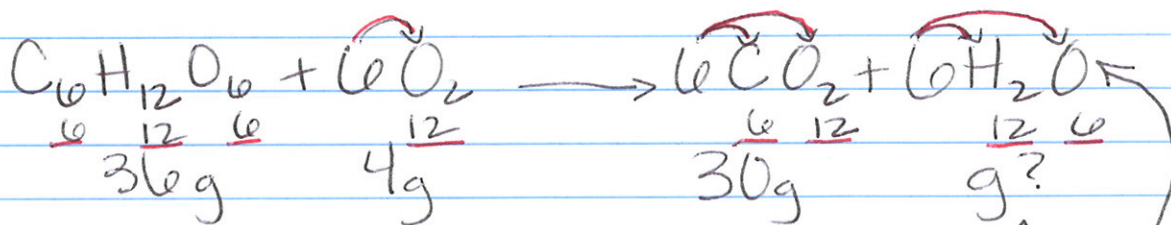


11:15 -  
11:48

Answer questions 1-6  
(see Atomic Structure questions 1-10)

11:50



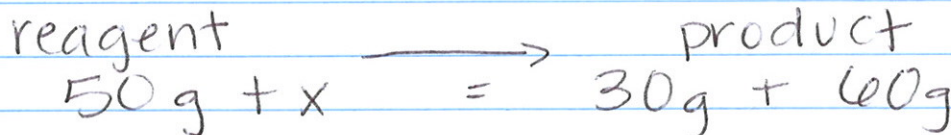
Mass is conserved

reagents

products

\* must be 10 because mass is conserved,  
 $36 + 4 = 40$        $30 + x = 40$        $x = 10$

solve



12:00

\* study for Test

Answering questions on Periodic Table elements

TEST TOMORROW

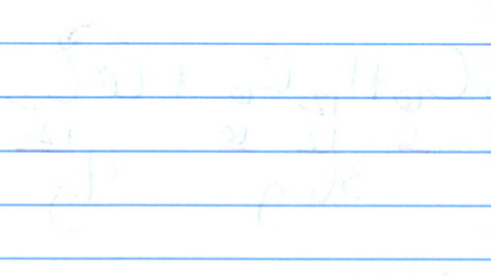
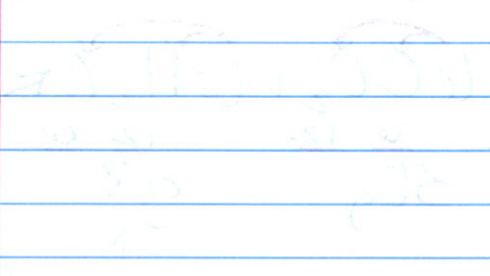
\* look over practice test

\* know highlighted elements

1/18/11  
1/19/11  
1/20/11

1/18/11  
1/19/11  
1/20/11

1/18/11  
1/19/11  
1/20/11



1/18/11  
1/19/11  
1/20/11

1/18/11  
1/19/11  
1/20/11

1/18/11  
1/19/11  
1/20/11

TEST 1 (1/18/11)

\* know probability distributions  
\* look over practice test

## Chemistry - Atomic Structure

10/3/11

1) the 3 particles of the Atom Are:

A \_\_\_\_\_

B \_\_\_\_\_

C \_\_\_\_\_

Their respective charges are

A

B

C

2) The number of protons in one atom of an element determines the atomic \_\_\_\_\_, and the number of electrons determines \_\_\_\_\_ of an element.

3) The atomic number tells you the number of \_\_\_\_\_ in one atom of an element. It also tells you the number of \_\_\_\_\_ in a neutral atom of that element.

The atomic number gives ~~you~~ the "identity" of an element as well as its location on the periodic table. No two elements will have the \_\_\_\_\_ atomic number.

4) The \_\_\_\_\_ of an element is the average mass of an element's naturally occurring atom, or isotopes, taking into account the \_\_\_\_\_ of each isotope.

5) The \_\_\_\_\_ of an element is the total number of protons & neutrons in the \_\_\_\_\_ of the atom.

6) The mass number is used to calculate the number of \_\_\_\_\_ in one atom of an element. In order to calculate the number of neutrons you must subtract the \_\_\_\_\_ from the \_\_\_\_\_.

7) Give the symbol & # of protons in one atom of:

Lithium \_\_\_\_\_ Bromine \_\_\_\_\_

Iron \_\_\_\_\_ Copper \_\_\_\_\_

Oxygen \_\_\_\_\_ Mercury \_\_\_\_\_

Krypton \_\_\_\_\_ Helium \_\_\_\_\_

8) Give the symbol & # of electrons in a neutral atom of:

Uranium \_\_\_\_\_ Chlorine \_\_\_\_\_

Boron \_\_\_\_\_ Iodine \_\_\_\_\_

Antimony \_\_\_\_\_ Xenon \_\_\_\_\_

9) Give the symbol & # of neutrons in one atom of:

\* To get "mass #" you must round the "atomic mass" to the nearest whole #. Show your calculations.

Barium \_\_\_\_\_ Bismuth \_\_\_\_\_

Carbon \_\_\_\_\_ Hydrogen \_\_\_\_\_

Fluorine \_\_\_\_\_ Magnesium \_\_\_\_\_

Europium \_\_\_\_\_ Mercury \_\_\_\_\_

10) Name the element which has the following number of particles:

a. 26 electrons, 29 neutrons, 26 protons \_\_\_\_\_

b. 53 protons, 74 neutrons \_\_\_\_\_

c. 2 electrons (neutral atoms) \_\_\_\_\_

d. 20 protons \_\_\_\_\_

e. 86 electrons, 125 neutrons, 82 protons (charged atom) \_\_\_\_\_

f. 0 neutrons \_\_\_\_\_