Cornell Notes

Name: Adriana Lopez / Tam

Date: 12/15/11

Topic:

Class/Subject: Chemistry p.4

11:20

chem. Rxns WS 1 →

#2

$2\,HBr \rightarrow \_\,H_2 + \_\,Br_2$

11:45 for lab:

	$NaPO_4$	Na_2SO_4	NaOH	NaCl
$Pb(NO_3)_2$	5 drops	↓	↓	↓
$Fe(NO_3)_3$	5 drops			
$Cu(NO_3)_2$	5 drops			

12:05 Questions Typed in lab report... Check Ms. Tam's website.

Due Wednesday →

In the Lab Today:

- go over Worksheet from yesterday

Single-replacement

$A + Bx \rightarrow Ax + B$

$Al + H_2SO_4 \rightarrow AlSO_4 + H_2$

(+3 −2)

$\rightarrow Al_2(SO_4)_3 + H_2$

*double-replacements WILL NOT be on Quiz

Go to stations & work on Labs

Questions

1) How would you describe a double replacement rxn to a friend who hasn't taken Chem?
2) What is a precipitate?
3) What indications did you observe today that suggest a chemical rxn happened?
4) Which ion (Pb^{2+}, Fe^{3+}, or Cu^{2+}) reacted with the most solutions?

→ BUT extra credit if turned in Tuesday!

*Quiz Tomorrow on Worksheet (chem. Rxns WS 1) on second page