

# Cornell Notes

Name Adriana Lopez / Tam

Date 12/15/11

Topic

Class/  
Subject Chemistry p.4

11:20  
chem. Rxns WS 1 →  
#2  
 $2\text{HBr} \rightarrow \text{H}_2 + \text{Br}_2$

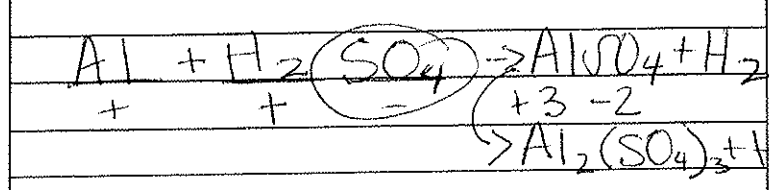
11:45 for lab:

	$\text{Na}_3\text{PO}_4$	$\text{Na}_2\text{SO}_4$	$\text{NaOH}$	$\text{NaCl}$
$\text{Pb}(\text{NO}_3)_2$ 5 drops	↓	↓	↓	↓
$\text{Fe}(\text{NO}_3)_3$ 5 drops	↓	↓	↓	↓
$\text{Cu}(\text{NO}_3)_2$ 5 drops	↓	↓	↓	↓

12:05 Questions  
Typed in lab report...  
Check Ms. Tam's website.  
Due Wednesday

In the Lab Today:  
• go over worksheet from yesterday

Single-replacement  
 $\text{A} + \text{Bx} \rightarrow \text{Ax} + \text{B}$



\*double-replacements  
WILL NOT be on Quiz

Go to stations & work on labs

- Questions
- 1) How would you describe a double replacement rxn to a friend who hasn't taken Chem?
  - 2) What is a precipitate?
  - 3) What indications did you observe today that suggest a chemical rxn happened?
  - 4) Which ion ( $\text{Pb}^{2+}$ ,  $\text{Fe}^{3+}$ , or  $\text{Cu}^{2+}$ ) reacted with the most solutions?

BUT extra credit if turned in Tuesday!

\*Quiz Tomorrow on Worksheet (chem. Rxns WS 1) on second page